

Freddies Dam Good Fries

G Psacharopoulos

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Introduction

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atoms change oxidation numbers
oxidation occurs when an atom's
oxidation state becomes more positive
indicating that electrons have been lost
reduction occurs when an atom's
oxidation state becomes less positive
indicating that electrons have been
gained

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commonly known as redox reactions
are reactions that involve the transfer
of electrons from one species to
another the species that loses electrons
is said to be oxidized while the species
that gains electrons is said to be

reduced

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of electron transfer recall that oxidation
involves loss of electrons and reduction
involves gain of electrons represent
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first one is done for you $23 \text{ Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + 2\text{Fe}$
 $\text{Fe} \rightarrow \text{Fe}^{2+} + 2\text{e}^{-}$ oxidation $\text{Co}^{2+} + 2\text{e}^{-} \rightarrow \text{Co}$ reduction $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$
 $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3\text{Ag}$ oxidation

reduction $25 \text{ Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation
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reactions indicate the elements that
undergo changes in oxidation number
and the magnitude of change then
identify the element oxidized and
reduced as well as the oxidizing agent
and reducing agent a $\text{i}_2 + 2\text{O}_5 + 5\text{Co} \rightarrow \text{i}_2 + 5\text{Co}_2 + 2\text{O}_5$

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is a transfer of electrons from one
substance to another are known as
oxidation reduction reactions or redox

reactions in this unit you will examine the oxidation reduction process and use the oxidation state and oxidation number concepts to identify redox reactions and to keep track of electrons

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and the reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$
 Sr^0 to Sr^{+2} oxidized reducing agent O^0 to O^{-2} reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ Li^0 to Li^{+1} oxidized red ag S^0 to S^{+2} reduced ox ag $3\text{Cs} + \text{Br}_2$