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web $m \times c \times \Delta t$ equation 1 where m is the total mass of the solution solute plus solvent c is the specific heat of the solution and Δt is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of

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web the heat of reaction ΔH_{soln} is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$

$\text{OH}^- \text{ (aq)} + \Delta H_{\text{soln}} 44.2 \text{ kJ mol rxn}$ write equations to similarly represent the dissolving process for

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and produces a large temperature
change in the form of heat in this
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solution will be calculated from the
data for six salts in order to

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filled with sodium acetate click the
metal disc and you will notice the liquid
changes state and becomes a solid in
the process heat is released the sodium
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other small piece of metal plastic bags
seal the coin inside a small bag place
this bag inside a larger bag containing

supersaturated sodium acetate get this
by dissolving as much sodium acetate
as possible in hot water this chemical
hand warmer is reusable

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research is to evaluate the various
number of compounds to be used as
hand warmers and locate the ones that
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use chemistry to determine which ionic
compound makes the best hand warmer
in this experiment you will collect data
that will allow you to calculate the heat
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of the substance being dissolved
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hand warmer is struck in a manner that
ruptures the inner pouch releasing the
ionic salt into the water of the outer

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