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Jicheng Xie

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in fact the general reaction between an acid and a base is ce
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pairs the reactants and the products

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and bases are weak that is they do not ionize fully in aqueous
solution a solution of a weak acid in water is a mixture of the
nonionized acid hydronium ion and the conjugate base of the
acid with the nonionized acid present in the greatest
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proton from a proton donor acid to a proton acceptor base in
a subsequent chapter of this text we will introduce the most
general model of acid base behavior introduced by the
american chemist g n lewis

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found in some common substances assume the solutions are
0 1 m at room temperature 2 at what ph or range of ph
values is a solution considered strongly acidic mildly acidic
neutral mildly basic and strongly basic provide a brief
explanation 3

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states that an acid is a substance that forms hydrogen ions H^+ when dissolved in water and a base is a substance that forms hydroxide ions OH^- when dissolved in water

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acids of CH_3NH_2 and S^{2-} are $CH_4NH_3^+$ and HS^- respectively
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