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web among the lanthanides in addition to 3 states 2 states are shown by nd sm

eu tm and yb only whereas 4 states is exhibited by ce pr nd tb and dy elements rest five elements show only 3 states oxidation states of lanthanide elements it has been shown that the lanthanide elements are highly electropositive and form essentially ionic

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symbol ln atomic number 57 cerium
 symbol ce atomic number 58
 praseodymium symbol pr atomic
 number 59 neodymium symbol nd
 atomic number 60 promethium symbol
 pm

the lanthanide elements z 58 71

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 comprise the largest naturally
 occurring group in the periodic table
 there are more than 100 minerals
 known to contain lanthanides but the
 only two of commercial importance are
 monazite and bastnaesite the chapter
 discusses the production of lanthanide
 elements by electrolysis of fused salts
 and by metallothermic reduction

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web the fifteen elements shown in table
 pageindex 1 from lanthanum la 4f 0 to
 lutetium lu 4f 14 are lanthanoids ln may
 be used as a general symbol for the
 lanthanoid elements although

lanthanoids scandium sc and yttrium y
 are sometimes called rare earth
 elements they are relatively abundant
 in the earth s crust

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 lanthanides has shown a continuous
 increase in recent years 50 taking all
 these facts into account it is reasonable
 to consider the development of
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 family of compounds at first the
 lanthanide elements principally cerium
 were used in protection against
 corrosion at high

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 essential biological role of rare earth
 elements lay hidden until the discovery
 in 2011 that lanthanides are specifically
 incorporated into a bacterial methanol
 dehydrogenase

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 displacement chromatography and
 elution chromatography as the two
 principal methods for separating
 lanthanide cations by ion exchange
 chromatography the ions to be
 separated are first sorbed onto a col
 umn of the resin the analyte metals are
 eluted from the column

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 located on the top row of the two rows
 of elements found below the main body
 of the periodic table placing the
 lanthanides and the actinides the row
 below the lanthanides below the other
 elements is just a matter of aesthetics
 the lanthanides belong to period 6 like
 cesium barium and hafnium