

Oh Ms Dura Que Mrmol A Mis Quejas

D Keegan

What makes for a good leaving group? + Example - Socratic Jun 23, 2015 · A good leaving group has to be able to part with its electrons easily enough, so typically, it must be a strong acid or weak base relative to other substituents on the same molecule. It helps to know the pKa of what would be leaving. Let's say you had a mechanism where you are trying to do an E2 reaction to make an -OH (hydroxyl) group leave. Maybe you have this compound on hand, sec-butanol ... **What is the formula of "magnesium hydroxide"? | Socratic Feb 27, 2017** · Mg (OH)₂ Now we know that hydroxides are salts of HO⁻, and some metal ion. Now if the parent metal has an electronic configuration of 2:8:2, then there are 12 electrons, and the atomic number of the metal is equal to 12. We look on the Periodic Table, and we find that Z=12, for "magnesium metal". As a Group 2 metal, magnesium forms a Mg⁽²⁺⁾ ion, and hence its hydroxide is Mg (OH)₂. *What is the chemical formula for lithium hydroxide? - Socratic Apr 15, 2017* · Lithium is a Group 1 metal and commonly forms a M⁺ ion. Hydroxide anion, -OH, has a unit negative charge. When they make music together, there is thus 1:1 stoichiometry between ions: Li(s) + H₂O(l) → LiOH (aq) + 1/2 H₂(g) ↑ □ □ □ **Question #44ec4 - Socratic** That's a silly way to say carbon dioxide! CO₂(g)+Ca(OH)₂(aq) rightleftharpoons Ca⁽²⁺⁾(aq)+H₂CO₃(aq) May be wrong, but it seems reasonable. The gas would interact with the hydroxide and water molecules to form carbonic acid. There are also concentrations of OH⁽⁻⁾ and H⁽⁺⁾ leftover from this reaction and from the auto-ionization of water, but that's probably going too far into detail. *0.10 moles of sodium cyanate dissolved in 250cm³ distilled Dec 28, 2015* · The cyanate anions will react with water to form isocyanic acid, HNCO, and hydroxide anions, OH⁻. This tells you that you can expect the pH of the resulting solution to be higher than 7. Calculate the molarity of the sodium cyanate solution by using the given number of moles and volume. Keep in mind that you need to use liters of solution and ... Could a buffered solution be made by mixing aqueous ... - Socratic Mar 25, 2017 · Could a buffered solution be made by mixing aqueous solutions of HCl and NaOH? Explain. Why isn't a mixture of a strong acid and its conjugate base considered a buffered solution? **What is the product of the following reaction? 1)CH₃OH - Socratic Dec 23, 2017** · These are ostensibly acid-base reactions.... For ammonium we could write.... NH₄⁺ + HO⁽⁻⁾ ⇌ NH₃(aq) + H₂O(l) For methanol, the acid base reaction would proceed ... **Question #93f99 - Socratic** The molarity of the acid is 1.434. Start with the equation: 2HNO₃ (aq)+Ca (OH)₂ (aq)⇌Ca (NO₃)₂ (aq))+2H₂O (l) This tells us that 2 moles of HNO₃ ... *What are the spectator ions in the reaction between sodium Oct 16, 2016* · The nitrate and the

sodium ions. $\text{Na}_2\text{CO}_3(\text{aq}) + 2\text{AgNO}_3(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s}) + 2\text{NaNO}_3(\text{aq})$ The net ionic equation is: $2\text{Ag}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s})$... **In each of the following pairs, which one has the lower viscosity?** Jul 31, 2015 · $\text{C}_2\text{H}_5\text{Cl}$ has weak dipole-dipole forces from the C-Cl bonds, but $\text{CH}_3\text{-OH}$ has strong hydrogen bonds. Confirmation: The viscosity of $\text{C}_2\text{H}_5\text{Cl}$ is 0.279 mPa·s, while the viscosity of $\text{CH}_3\text{-OH}$ is 0.593 mPa·s.

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