

# The Friend Zone Abby Jimenez

Robert R. Redfield

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web the relative electronegativities of the bonded atoms is h c cl and so the

bond moments all point toward the cl end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as bf 3 trigonal planar ch 4 tetrahedral pf 5 trigonal bipyramidal and sf 6

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web explain the difference between a nonpolar covalent bond a polar covalent bond and an ionic bond answer nonpolar covalent electronegativity difference is less than 0 4 nonmetal nonmetal close together on the periodic table

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web apr 1 2021 updated on april 01 2021 chemical bonds may be classified as being either polar or nonpolar the difference is how the electrons in the bond are arranged key takeaways what is a polar bond in chemistry a polar bond is a type of covalent bond in which the electrons forming the bond are unequally distributed