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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 4 Al + 3 $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield yield actual $\times 100$ 260.0 g $\times 100 / 86.01$ theoretical 302.3 g

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web limiting reagents and percentage yield worksheet 1 consider the reaction $\text{i}_2\text{o}_5 + \text{g }_5\text{co} \rightarrow \text{g }_5\text{co}_2 + \text{g }_2\text{i}$ g a 80 0 grams of iodine v oxide i_2o_5 reacts with 28 0 grams of carbon monoxide co determine the mass of iodine i_2 which could be produced 80 g i_2o_5 1 mol i_2o_5 1 mol i_2 xs 1 333 8 g i_2o_5 1 mol i_2o_5 28 g co 1 mol co

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3