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 equation below 4 no g 6 h 2 o l 4 nh 3 g
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 endothermic rh gt 0 b how much heat
 will be absorbed or released if 10 g of
 no is reacted with 10 g h 2 o

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 reaction was multiplied by 2 therefore
 Δh must also be halved 3 answer c
 explanation option c is incorrect
 because 566 kj mol is the Δh per mole
 of o2 and not for 2 moles even though
 options b and d are expressed in grams
 they are still correct because 56

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 change when 472 g of carbon reacts

with excess O_2 according to the following equation how much heat energy must be expelled by the body to rid the body of one pound of fat $2s\ 3o\ 2\ 2so\ 3\ h\ 7914\ kJ$ is this endothermic or exothermic 2

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web $q\ 140\ g\ 4\ 184\ j\ 16oc\ g\ oc\ q\ 9372\ 16\ j\ 35\ 5\ g\ x\ 1\ mole\ kcn\ 0\ 5453\ moles$

$65\ 1\ g\ 0\ 5453\ mole\ moles\ 5$ when a $15\ g$ sample of $srcl_2$ dissolves in $80\ g$ of water in a calorimeter the temperature drops from $42\ oc$ to $15\ c$ calculate the Δh for the process

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