

Oh The Places You'll Go Quotes Dr Seuss

Arturo Cuomo

Question #0a25a - Socratic Mar 20, 2017 · Lithium hydroxide has a mass of 24 g per mole, so 2.50 g is 0.104 mole. Dissolve this in 1.50 L, and the concentration is $[\text{LiOH}] = 0.104 \text{ mol} / 1.50 \text{ L} = 0.0694 \text{ mol/L}$. Since LiOH is a strong base, it ionizes completely $\text{LiOH} \rightarrow \text{Li}^+ + \text{OH}^-$. Since each one mole of LiOH produces one mole of OH^- , the concentration of OH^- will be 0.0694 mol/L as well. *What is the chemical formula for lithium hydroxide?* - Socratic Apr 15, 2017 · Lithium is a Group 1 metal and commonly forms a M^+ ion. Hydroxide anion, OH^- , has a unit negative charge. When they make music together, there is thus 1:1 stoichiometry between ions: $\text{Li(s)} + \text{H}_2\text{O(l)} \rightarrow \text{LiOH(aq)} + \text{H}_2\text{(g)}$ ↑ □ □ □ *Can you give the IUPAC name for the following (CH_3)₃C-OH ...* Aug 29, 2016 · For tert-butyl alcohol, $(\text{CH}_3)_3\text{C}-\text{OH}$, clearly the longest chain is three carbons long, and C2 is substituted by OH , and CH_3 . So this is a propanol derivative: 2-methylpropan-2-ol For isopropyl alcohol, $\text{H}_3\text{C}-\text{CH}(\text{OH})\text{CH}_3$, the longest chain is again three carbons long, and C2 is substituted by OH , so propan-2-ol I think this is right, and I ... **Question #e1a77 - Socratic** Answer is 286g (3s.f) Concept required: mole calculation First start off by finding the number of moles for both compounds: $\text{PbCl(OH)} = 0.185 / (207.2 + 35.5 + 16 + 1) \times 1000 \text{ (1kg=1000g)} = 0.712 \text{ (to 3s.f)}$ $\text{Pb}_2\text{Cl}_2\text{CO}_3 = 0.185 / (207.2 \times 2 + 35.5 \times 2 + 12 + 16 \times 3) \times 1000 = 0.339 \text{ (to 3s.f)}$ Next, find the moles of Pb from each compound: Pb : PbCl(OH) 1 : 1 Therefore, moles of Pb in PbCl(OH) is 0.712 Pb : $\text{Pb}_2\text{Cl}_2\text{CO}_3$ 2 : 1 ... Question #bcb2d - Socratic calculate moles of Acid, calculate moles of Base. subtract, $\#\text{OH}^-$ -# should be in excess, calculate Molarity of $\#\text{OH}^-$ -, $\text{pOH} = -\log(\#\text{OH}^-)$, then $14 - \text{pOH} = \text{pH}$ What is the product of the following reaction? 1) CH_3OH - Socratic Dec 23, 2017 · These are ostensibly acid-base reactions.... For ammonium we could write.... $\text{NH}_4^+ + \text{HO}^- \rightarrow \text{NH}_3(\text{aq}) + \text{H}_2\text{O(l)}$ For methanol, the acid base reaction would proceed ... **In each of the following pairs, which one has the lower viscosity?** Jul 31, 2015 · $\text{C}_2\text{H}_5\text{Cl}$ has weak dipole-dipole forces from the C-Cl bonds, but $\text{CH}_3\text{-OH}$ has strong hydrogen bonds. Confirmation: The viscosity of $\text{C}_2\text{H}_5\text{Cl}$ is 0.279 mPa·s, while the viscosity of $\text{CH}_3\text{-OH}$ is 0.593 mPa·s. *What reaction represents formation of Cu(OH)_2 ?* | Socratic Copper forms an insoluble hydroxide that is fairly poorly characterized; another way we could look at it as a hydrated copper oxide, i.e. $\text{CuO} \cdot \text{H}_2\text{O} \equiv \text{Cu(OH)}_2$. **Question #d6b18 - Socratic** We want the standard enthalpy of formation for Ca(OH)_2 . Thus, our required equation is the equation where all the constituent elements combine to form the compound, i.e.: $\text{Ca} + \text{H}_2 + \text{O}_2 \rightarrow \text{Ca(OH)}_2$ Let us now write down the given equations: [The first equation mentioned is incorrect, and so I have revised it.] (1) $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O(l)}$ and $\Delta H_1 = -571.66$

kJmol^{-1} (2) $\text{CaO (s)} + \text{H}_2\text{O (l)}$... *Calculating the concentration of excess HCL in E - Socratic* The acid in excess is then titrated with NaOH (aq) of KNOWN concentration....we can thus get back to the concentration or molar quantity of M (OH)_2 ...as it stands the question (and answer) are hypothetical...

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