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2 1 Be NO_3 2 2 RbF

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2 H_2O 8 1 C_3H_8 5 O_2 3 CO_2 4

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2 5 Cl_2 KBr KCl Br_2 6 N_2 H_2 NH_3 7 NdF_3

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5 OH

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for balancing chemical equations is to make both sides of the reaction the reactants and the products equal in the number of atoms per element this stems from the universal law of the conservation of mass which states that matter can neither be created nor destroyed

**answer key for the balance
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equation $\text{mg} + \text{oh}^- + 2 \text{hcl} + \text{mgcl}_2 + \text{h}_2\text{o}$ note
all reactants and products require a
coefficient of at least one stuck review
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