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 curve is a plot of the ph of a solution as  
 a function of the amount of base usually  
 in ml added to an acidic solution or as  
 the amount of acid added to a boss

solution the base is added to

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web mar 7 2021 unlike strong acids or bases the shape of the titration curve for a weak acid or base depends on the  $pK_a$  or  $pK_b$  of the weak acid or base being titrated a solution  $pH$  as a function of the volume of 1.00 M  $NaOH$  added to 10.00 mL of 1.00 M solutions of weak acids with the indicated  $pK_a$  values

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feedback objective to determine the

amount of substance in a solution of

unknown concentration using various

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word titration comes from the latin

word titulus which means inscription or title

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base of the acid the four parts of the

titration curve are described below and

you should look to the appropriate text

section to see how they are treated

pure acid 0 mL of base is added section

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web solution a titrant volume 0 mL the

solution  $pH$  is due to the acid ionization

of  $HCl$  because this is a strong acid the

ionization is complete and the

hydronium ion molarity is 0.100 M the

$pH$  of the solution is then  $pH = \log 0.100 = 1.000$

$pH = \log 0.100 = 1.000$  b titrant volume

12.50 mL

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analytical procedure used to determine

the accurate concentration of a sample

by reacting it with a standard solution

one type of titration uses a

neutralization reaction in which an acid

and a base react to produce a salt and

water

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web at the equivalence point in an acid

base titration moles of base moles of

acid and the solution only contains salt

and water acid base titrations are

monitored by the change of  $pH$  as

titration progresses

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example is based upon henderson

hasselbalch equation the equation gives

relation between  $pH$  and  $pH_a$  of weak

acid the equation is as  $pH = pK_a + \log \frac{[salt]}{[acid]}$

at equivalence point all a view the

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of the weak acid or base being

titrated a solution  $pH$  as a function of

the volume of 1.00 M  $NaOH$  added to 10

00 mL of 1.00 M solutions of weak acids

with the indicated  $pK_a$  values

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chemistry atoms first 2e strong acid

strong base titrations when titrating a

strong acid such as hydrochloric acid

with sodium hydroxide you are reacting

the  $HCl$  directly with  $NaOH$   $HCl(aq) + NaOH(aq) \rightarrow H_2O(l) + NaCl(aq)$

$HCl(aq) + NaOH(aq) \rightarrow H_2O(l) + NaCl(aq)$

nacl a q

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purpose the purpose of this lab is to determine the shape of acid base titration curves theoretically and experimentally and to use an acid base titration to measure the effectiveness of antacid tablets summary a titration curve is the plot of the ph value of a solution vs the amount of titrant added *acid base titrations chemistry libretexts* - Feb 28 2023

web aug 30 2022 acid base titrations are usually used to find the amount of a known acidic or basic substance through acid base reactions the analyte titrand is the solution with an unknown molarity the reagent titrant is the solution with a known molarity that will react with the analyte

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or 2 below 3 estimate of the midpoint of the steeply rising section of the titration curve provides an evaluation of the equivalence point

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jessica chung chem 267 lab date

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base titration curve from titration with

indicator for part 1 volume of naoh

added ml ph of the solution 4 5 4 6 4 6

5 9 5 9

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calculations and titrimetric analysis the volume of a solution required to react

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web background plot of the ph of a solution against the volume of titrant added is called a titration curve the ph can be measured directly with a ph meter while titrant is added from a buret for the acids used in this experiment the titrant will always be a

0 1 m solution of the strong base naoh

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started with a slow gradual change in ph as it reached the equivalence point  
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