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 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) \rightarrow \frac{7}{2} \text{O}_2(\text{g}) + 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$
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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ}$
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web this activity provides a demonstration of hess law using three reactions the solubility NaOH in water the solubility NaOH in HCl and the reaction of a solution of HCl and a solution of NaOH

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web hess s law worksheet answers 1 calculate h for the reaction $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$ from the following data $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$ $\Delta H = 1411 \text{ kJ}$ $\text{C}_2\text{H}_6 + \frac{3}{2}\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$ $\Delta H = 1560 \text{ kJ}$ *7a first law enthalpy calorimetry and hess s law worksheet* - Oct 05 2022 web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one

body to another in a sense work is energy in the process of transfer

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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kj in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and Δh values to determine the heat of reaction for the reaction c 2 h 4

experiment 9 enthalpy of reaction

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web thus for reaction 9 2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O h 55 8 kj as

shown note that the enthalpy change ΔH products reactants is positive if heat is absorbed that is if ΔH products

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web hess s law enthalpy of reaction and hess s law abstract the goal of this lab is to perform experiments and practice hess s law by determining enthalpy change in chemical reactions in this lab three experiments were conducted in order to test hess s law the experiments required observing and collecting data on three different reactions

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{aq} + \text{OH}^- \text{aq}$ h 1 2 $\text{NaOH (s)} \rightarrow \text{H}^+ \text{aq} + \text{Cl}^- \text{aq}$ h 2 3 $\text{O}^- \text{Na}^+ \text{aq} \rightarrow \text{Cl}^- \text{aq}$ h 2 3