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the enthalpies free energy of the products and the reagents  
keeping into account the

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web answer a 1 explanation the entire reaction was  
multiplied by 2 therefore  $\Delta h$  must also be halved 3 answer c  
explanation option c is incorrect because 566 kj mol is the  $\Delta h$   
per mole of o2 and not for 2 moles even though options b and

d are expressed in grams they are still correct because 56  
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 web thermochemistry practice worksheet answer key what is  
 the heat change when 472 g of carbon reacts with excess  $O_2$   
 according to the following equation how much heat energy  
 must be expelled by the body to rid the body of one pound of  
 fat  $2s\ 3o\ 2\ 2so\ 3h\ 7914kj$  is this endothermic or exothermic  
 2

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 thermal energy in chemical reactions exothermic transfer of  
 heat from the system to the surroundings  $2h\ 2\ g\ o\ 2\ g\ 2h\ 2\ o$   
 $l$  energy endothermic the transfer of heat from the  
 surroundings to the system energy  $2hgo\ s\ 2hg\ l\ o\ 2\ g$   
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 that show the associated enthalpy change enthalpy is an  
 extensive property depends on amount of matter present this  
 fact means that the magnitude of  $h$  is directly proportional to  
 the amount of reactant consumed in the process  
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 web  $q\ 140\ g\ 4\ 184\ j\ 16oc\ g\ oc\ q\ 9372\ 16\ j\ 35\ 5\ g\ x\ 1\ mole$   
 $kcn\ 0\ 5453\ moles\ 65\ 1\ g\ 0\ 5453\ mole\ moles\ 5$  when a 15 g  
 sample of  $srcl_2$  dissolves in 80 g of water in a calorimeter the  
 temperature drops from 42  $oc$  to 15  $c$  calculate the  $\Delta h$  for

the process  
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 web thermochemistry answers and explanations 1 a a  
 thermodynamic system has two means of exchanging energy  
 with its surroundings heat flow and pressure volume work  
 heat flow  $q$  is the transfer of thermal energy due to  
 temperature difference work  $w$  is macroscopic energy  
 transfer between the system and its surroundings that  
 occurs

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web in this set of practice questions we will summarize the  
 main concepts of thermochemistry such as the relationship  
 between internal energy work and heat exothermic and  
 endothermic process heat capacity constant pressure  
 calorimetry constant volume calorimetry the enthalpy the  
 standard enthalpies read more

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web jan 10 2021  $\Delta h\ 27\ 4\ kj\ mol$  and  $\Delta s\ 29\ 8\ j\ k$  sulfur  
 trioxide  $so_3$  is produced in large quantities in the industrial  
 synthesis of sulfuric acid sulfur dioxide is converted to sulfur  
 trioxide by reaction with oxygen gas write a balanced  
 chemical equation for the reaction of  $so_2$  with  $o_2\ g$  and  
 determine its  $\Delta g$