

# The Sun Is Up The Sky Is Blue

Michael Rajnik

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key 1 find the  $\Delta h$  for the reaction below given the following reactions and subsequent  $\Delta h$  values  
 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g})$   $\Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g})$   $\Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2(\text{g})$   $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$  from the following data  
 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$   $\Delta h = 1411 \text{ kJ}$  mole  
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$   $\Delta h = 1560 \text{ kJ}$  mole  
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 285 \text{ kJ}$  mole

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 $\text{Mg}(\text{s}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{MgO}(\text{s})$  we used *additivity of heats of reaction hess s law lab report* - Aug 03 2022  
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cohen 2016 the equations for the reactions in the experiment done are as follows  
 $\text{NaOH}(\text{s}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{NaOH}(\text{aq})$   
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reactions the  $\Delta h$  for the overall process must be the sum of the  $\Delta h$  values of the constituent reactions enthalpy change  $\Delta h$  is independent of the path that a reaction follows to move from reactants

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work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kJ in your final answer multiply the mass by the change in temperature and the  $c$  given trial 1 use hess s law and the following equations and  $\Delta h$  values to determine the heat of reaction for the reaction  $\text{C}_2\text{H}_4$

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{O}$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta h = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta h_{\text{products}}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions  $\text{NaOH}(\text{s}) + \text{Na}(\text{aq}) \rightarrow \text{OH}(\text{aq})$   $\text{H}^+(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_3\text{O}^+(\text{aq})$   $\text{Cl}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{HCl}(\text{aq})$