

Atlas Del Jazz En Mexico Pdf

Roman Wölfel

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approaches to capture its audience s
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bases its appeal on provoking strong
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so happy to be here and i want to thank
audrey beverly and the rest of the west
valley relay committee for inviting me
this relay was the first i ever
experienced two years ago i came to
this relay as a survivor after 8 hard
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commonly known as redox reactions
are reactions that involve the transfer
of electrons from one species to
another the species that loses electrons
is said to be oxidized while the species
that gains electrons is said to be
reduced

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of electron transfer recall that oxidation
involves loss of electrons and reduction
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redox reaction this is a reaction in
which electrons are transferred

between reactants what do the
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reduction reduction is a gain of
electrons what is a helpful way to
remember this
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reactions for each redox reaction the
first one is done for you 23 $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ $\text{Co}^{2+} + \text{e}^- \rightarrow \text{Co}$
24 3 $\text{Ag} \rightarrow \text{Ag}^+ + \text{e}^-$ $\text{Ni}^{2+} + 2\text{e}^- \rightarrow \text{Ni}$
25 $\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}$ $\text{Pb} \rightarrow \text{Pb}^{2+} + 2\text{e}^-$
26 $\text{O}_2 + 4\text{e}^- \rightarrow 2\text{O}^{2-}$

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web in each of the following redox

reactions indicate the elements that
undergo changes in oxidation number
and the magnitude of change then
identify the element oxidized and
reduced as well as the oxidizing agent
and reducing agent a $\text{I}_2 + 5\text{CO} \rightarrow 2\text{IO}_3^- + 5\text{C}_2\text{O}_4^{2-}$

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web chemical reactions in which there
is a transfer of electrons from one
substance to another are known as
oxidation reduction reactions or redox
reactions in this unit you will examine
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number concepts to identify redox
reactions and to keep track of electrons
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what is oxidized and what in each
reaction identify the oxidizing agent
and the reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$
 $\text{Sr} \rightarrow \text{Sr}^{2+} + 2\text{e}^-$ $\text{O}_2 + 4\text{e}^- \rightarrow 2\text{O}^{2-}$
oxidized reducing agent O_2 to O^{2-}
reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$
oxidized red ag $\text{S} \rightarrow \text{S}^{2-}$ reduced ox ag
 $3\text{Cs} + \text{Br}_2$