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R Bogdan

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concentrations of the hydrogen ions

and hydroxide ions acidic solution one

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solution one in which the H^+ is less than

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basic solution is less than $1 \times 10^{-7} \text{ M}$

they are the same as basic solutions

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strongest base we can determine their

weakest conjugate acid the conjugate

acids of CH_3NH_2 and S^{2-} are CH_4NH_3^+

and HS^- respectively next we consider

that acidity increases with positive

charge on the molecule thus ruling out

that S^{2-} is the weakest base

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