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reaction $\text{ce zn cucl}_2 \text{ zncl}_2 \text{ cu zn cucl}_x \text{ 2}$
 $\text{zncl}_x \text{ 2 cu}$ how many moles of ce zncl_2
 $\text{zncl}_x \text{ 2}$ will be produced from 23.0 g of
ce zn zn assuming ce cucl_2
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 $\text{c}_6\text{h}_{10} \text{ 17 o}_2 \text{ 12 co}_2 \text{ 10 h}_2\text{o}$ br 3 1 hbr 1
 $\text{khco}_3 \text{ 1 h}_2\text{o} \text{ 1 kbr 1 co}_2 \text{ br 4 2 gabr}_3 \text{ 3}$
 $\text{na}_2\text{so}_3 \text{ 1 ga}_2 \text{ so}_3 \text{ 3 6 nabr br 5 3 sno 2}$
 $\text{nf}_3 \text{ 3 snf}_2 \text{ 1 n}_2\text{o}_3 \text{ br}$ using the following
equation $\text{br 2 naoh h}_2\text{so}_4 \text{ 2 h}_2\text{o na}_2\text{so}_4$
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 $\text{khco}_3 \text{ h}_2\text{o kbr co}_2 \text{ gabr na}_2\text{so}_3 \text{ gaz so}_3$
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3.00 kg of sic answer $\text{ce sio}_2 \text{ 3c}$
 $\rightarrow \text{sic 2co}$

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following equations $\text{n}_2 \text{ f}_2 \text{ nf c}_6\text{h}_{10} \text{ 02}$
 $\text{c}_0 \text{ ho 3 hbr khco}_3 \text{ h}_2\text{o kbr 4 gabry na}$
 $\text{so}_3 \text{ gaz so}_3 \text{ 3 5 sno nf snf}_2 \text{ n}_2\text{o co}_2$
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this is the mole ratio between two
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