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Emilie Sanchez

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point of an acid base titration is the
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completely with the other component
the equivalence point in the titration of
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ph 7 0

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partly because so many compounds can
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are those of a strong acid with a strong
base table 15 7 1 shows data for the
titration of a 25 0 ml sample of 0 100 m
hydrochloric acid with 0 100 m sodium
hydroxide the values of the ph
measured after successive additions of

small amounts of naoh are listed in the
first column of this table and are
graphed in

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reaction between an acid and a base
can be very useful if an acidic solution
of known concentration a standard
solution is added to a basic alkaline
solution of unknown concentration until
the solution is exactly neutralised i e
there is only salt and water it is
possible to calculate

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acid with a base or vice versa the
indicator is a substance that can exist
in two forms an

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neutralizations in which there is a 1:1 ratio between the acid and the base the example below demonstrates the technique to solve a titration problem for a titration of sulfuric acid with sodium hydroxide

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concentration the titrant is added to a solution of the substance being studied the analyte in an acid base titration the titrant is a strong base or a strong acid and the analyte is an acid or a base respectively

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