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 enthalpy changes $h_2O(s) \rightarrow h_2O(l)$ 6 01
 kJ mol this is an endothermic process it
 requires 6 01 kJ to melt one mole of ice
 $h_2O(s)$ the enthalpy value will change if
 the number of moles varies from the 1
 1 reaction stoichiometry this is an
 exothermic process
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 web in this set of practice questions we
 will summarize the main concepts of
 thermochemistry such as the
 relationship between internal energy
 work and heat exothermic and
 endothermic process heat capacity
 constant pressure calorimetry constant
 volume calorimetry the enthalpy the
 standard enthalpies of formation and
 their use in
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 web we will revisit this today and go on
 to see that if we use a special kind of
 thermochemical reaction called the
 standard enthalpy of formation we can
 calculate enthalpies of reactions

without having to manipulate a series
 of individual thermochemical equations
 for each step
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 web we can express the first law in
 many ways one of the more useful
 expressions is that the change in
 internal energy e of a system in any
 process is equal to the heat q added to
 the system plus the work w done on the
 system by its surroundings e
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