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web when the atoms in a bond are not the same the bonding electrons are shared and the bond is electronegativities equally unequally non polar polar unequally polar the degree of polarity of a bond between any two atoms is determined by consulting a table of

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web chemical bonds hold molecules together and create temporary connections that are essential to life types of chemical bonds including covalent ionic and hydrogen bonds and london dispersion forces introduction living things are made up of atoms but in most cases those atoms aren't just floating around individually  
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web explain how a molecule that

contains polar bonds can be nonpolar  
answer problem 6 2 2 6 2 2 which of the following molecules and ions contain polar bonds which of these molecules and ions have dipole moments a clf<sub>5</sub> b clo<sub>2</sub> c tecl<sub>2</sub> d tecl<sub>4</sub> e pcl<sub>3</sub> f sef<sub>4</sub> g ph<sub>3</sub> h ph<sub>2</sub> i xef<sub>2</sub>  
2 answer problem 6 2 3 6 2 3

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web the relative electronegativities of the bonded atoms is h c cl and so the bond moments all point toward the cl end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as bf<sub>3</sub> trigonal planar ch<sub>4</sub> tetrahedral pf<sub>5</sub> trigonal bipyramidal and sf<sub>6</sub>

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web explain the difference between a nonpolar covalent bond a polar covalent bond and an ionic bond  
answer nonpolar covalent electronegativity difference is less than 0.4 nonmetal nonmetal close together on the periodic table  
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2021 chemical bonds may be classified  
as being either polar or nonpolar the

difference is how the electrons in the  
bond are arranged key takeaways what  
is a polar bond in chemistry a polar

bond is a type of covalent bond in  
which the electrons forming the bond  
are unequally distributed