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 methanol CH_3OH can be produced by the following reaction
 $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68 g
 of CO is reacted with 8.6 g of H_2

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 3 sto 5 differentiate between the actual yield and theoretical
 yield of a chemical reaction

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 consider the following reaction $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{PO}_4$
 NaNO_3 which reactant is limiting assuming we started with
 30.0 grams of ammonium nitrate and 50.0 grams of sodium
 phosphate what is the

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 consider the reaction $\text{I}_2\text{O}_5 + \text{CO} \rightarrow \text{CO}_2 + \text{I}_2$ a 80.0 grams
 of iodine V oxide I_2O_5 reacts with 28.0 grams of carbon
 monoxide CO determine the mass of iodine I_2 which could be
 produced 80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333.8 g I_2O_5 1
 mol I_2O_5 28 g CO 1 mol CO

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the actual yield to the theoretical yield multiplied by 100 to
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