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 $\text{NaBr} + \text{CaOH}_2 \rightarrow \text{CaBr}_2 + \text{NaOH} + \text{NH}_3$
 $\text{H}_2\text{SO}_4 + \text{NH}_4_2\text{SO}_4 \rightarrow \text{C}_5\text{H}_9\text{O}_2 + \text{CO}_2 + \text{H}_2\text{O}$
 $\text{Pb} + \text{H}_3\text{PO}_4 \rightarrow \text{H}_2 + \text{Pb}_3\text{PO}_4$
 $2\text{Li}_3\text{N} + \text{NH}_4\text{NO}_3 \rightarrow \text{LiNO}_3 + \text{N}_2$

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$\text{OH}_2 + \text{Al}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{AlOH}_3$ type of reaction double displacement 3 $3\text{Mg} + \text{Fe}_2\text{O}_3 \rightarrow 3\text{MgO} + \text{Fe}$ type of reaction single displacement 4 $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$ type of reaction combustion 5

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