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main group to understand theory behind lewis structure and  
valence shell electron pair repulsion theory to experience  
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represented by the given option a in this structure the  
oxygen atom is in the center and is bonded to two hydrogen  
atoms the oxygen atom has two lone pairs of electrons this  
structure follows the octet rule where each atom has a full  
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compounds depends on the types of bonds that exist within  
the molecule and on the overall shape of the molecules the  
branch of chemistry that investigates how molecules are put  
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in the molecule as lone pairs unshared electrons drawn as  
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electrons that each atom contributes to the molecule  
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web the pcl 3 lewis structure has the typical case of  
phosphorus p in the center with 3 bonds to 3 other atoms  
phosphorus is from the same column as nitrogen in the  
periodic table meaning that p and n generally have the same  
bonding structure note the lone pair dots without bonds on  
top of p just like for n in the previous example for nh 3  
chlorine cl

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have based on the number of valence electrons in each atom  
when drawing the structure of an ion be sure to add subtract  
electrons to account for the charge step 2 connect the atoms  
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web this exercise required us to take the formulas in the data table and calculate the number of valence electrons using the periodic table draw its lewis structure on paper build the molecular model and record a photo list the number of bonds and lone pairs surrounding the central atom and lastly identify the structure geometry

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