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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml

bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract *experiment synthesis of 1 bromo 4 nitrobenzene pbworks* - May 28 2023 in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid

react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical x 100 what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for

this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 o  
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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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**nitration of bromobenzene 10 19 -**

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 over 5 minutes slowly pipet this  
 bromobenzene into the icy cold  
 nitration mixture hno 3 h 2 so 4 in the  
 125 ml flask and stir continue to stir  
 the reaction mixture in the ice bath for  
 30 minutes do not heat in a new 250 ml  
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 on the right the arenium cation has a  
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 substitution mechanism was used to  
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 where substituents are added on a  
 benzene ring by way of electrophilic  
 aromatic substitution the electrophilic  
 source in the nitration of  
 bromobenzene comes from a nitronium  
 ion no2 which acts as the electrophilic  
 species  
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 from 4 0ml of nitric acid and 4 0ml of  
 sulfuric acid and will then react with 4  
 5ml of bromobenzene this will produce  
 o bromonitrobenzene and p  
 bromonitrobenzene

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 allows for intermediate chemical  
 structures to be used as precursors to  
 pharmaceuticals and other industrially  
 produced chemicals by adding different  
 functional groups into an aromatic ring  
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for this experiment nitration of  
bromobenzene was the focus of the lab  
the benzene is an aromatic compound

and when it reacts with wither a  
mixture of sulfuric acid or nitric acid  
creates what is known as a nitro group