

# Oh I Believe In Yesterday Lyrics

JA Banks

What is the chemical formula for lithium hydroxide? - Socratic Apr 15, 2017 · Lithium is a Group 1 metal and commonly forms a  $M^+$  ion. Hydroxide anion,  $OH^-$ , has a unit negative charge. When they make music together, there is thus 1:1 stoichiometry between ions:  $Li(s) + H_2O(l) \rightarrow LiOH(aq) + \frac{1}{2} H_2(g)$  ↑ □ □ **What are the spectator ions in the reaction between sodium** Oct 16, 2016 · The nitrate and the sodium ions.  $Na_2CO_3(aq) + 2AgNO_3(aq) \rightarrow Ag_2CO_3(s) + 2NaNO_3(aq)$  The net ionic equation is:  $2Ag^+(aq) + CO_3^{2-}(aq) \rightarrow Ag_2CO_3(s)$  ... **Could a buffered solution be made by mixing aqueous ...** - Socratic Mar 25, 2017 · Could a buffered solution be made by mixing aqueous solutions of HCl and NaOH? Explain. Why isn't a mixture of a strong acid and its conjugate base considered a buffered solution? **What is the formula of "magnesium hydroxide"?** | Socratic Feb 27, 2017 ·  $Mg(OH)_2$  Now we know that hydroxides are salts of  $HO^-$ , and some metal ion. Now if the parent metal has an electronic configuration of 2:8:2, then there are 12 electrons, and the atomic number of the metal is equal to 12. We look on the Periodic Table, and we find that  $Z=12$ , for "magnesium metal". As a Group 2 metal, magnesium forms a  $Mg^{2+}$  ion, and hence its hydroxide is  $Mg(OH)_2$ . **If the pH=9 for a solution of KOH(aq), what is the ...** - Socratic Aug 18, 2016 · If the pH = 9 for a solution of KOH(aq), what is the concentration of  $OH^-$ ? *In each of the following pairs, which one has the lower viscosity?* Jul 31, 2015 ·  $C_2H_5Cl$  has weak dipole-dipole forces from the C-Cl bonds, but  $CH_3-OH$  has strong hydrogen bonds. Confirmation: The viscosity of  $C_2H_5Cl$  is 0.279 mPa·s, while the viscosity of  $CH_3-OH$  is 0.593 mPa·s. **What makes for a good leaving group? + Example** - Socratic Jun 23, 2015 · A good leaving group has to be able to part with its electrons easily enough, so typically, it must be a strong acid or weak base relative to other substituents on the same molecule. It helps to know the  $pK_a$  of what would be leaving. Let's say you had a mechanism where you are trying to do an  $E2$  reaction to make an  $-OH$  (hydroxyl) group leave. Maybe you have this compound on hand, sec-butanol ... **What is the product of the following reaction? 1)  $CH_3OH$**  - Socratic Dec 23, 2017 · These are ostensibly acid-base reactions.... For ammonium we could write....  $NH_4^+ + OH^- \rightarrow NH_3(aq) + H_2O(l)$  For methanol, the acid base reaction would proceed ... **Question #44ec4** - Socratic That's a silly way to say carbon dioxide!  $CO_2(g) + Ca(OH)_2(aq) \rightarrow Ca^{2+}(aq) + H_2CO_3(aq)$  May be wrong, but it seems reasonable. The gas would interact with the hydroxide and water molecules to form carbonic acid. There are also concentrations of  $OH^-$  and  $H^+$  leftover from this reaction and from the auto-ionization of water, but that's probably going too far into detail. *Question*

#627ef - Socratic functional groups like -OH, -CHO, -CO- are present in carbohydrates. commonly two functional are present in a carbohydrate, -OH and -CHO, or -OH and -CO-. Aldose shows different reactions and test like Tollen's silver mirror test, felling's test. different types of reactions are also shown by carbohydrates like Osazone synthesis, length increasing reaction, ...

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